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SOURCE Documentary as indicated. (Information requested.)

RECENTLY PUBLISHED RESEARCH OF
THE VORONEZH STATE UNIVERSITY

"Study of the Melting-Point Diagram in the System $KHS-KNO_2$," N. A. Palkina, Voronezh State U

"Zhar Obshch Khimii" Vol 15, 1945, pp 911-14

Diagrams of state were established for the KCNS-KNO₂ system. It was found that a compound KCNS-KNO₂ with a noncongruent melting point exists. A eutectic point was found at 120° and 66% KCNS and a transition point at 128° and 58% KCNS. On heating to 400° the compounds explode and the nature of the explosives indicates the fragmentation power of experimental mixtures of KCNS-KNO₂.

"Boron Fluoride as a Catalyst in Organic Chemistry: IV. Condensation of Phenol with Pseudobutylene," S. V. Zavgerodny, Vornesh State U

"Zhur Obshch Khimi" Vol 16, 1946, pp 1495-1504

PhOH, H_2SO_4 , and 2-butene were discharged into a salt-ice cooled autoclave and the mixture was agitated at the desired temperature, after which an extraction with 10% KOH, followed by the usual treatment, gave the phenolic products; the neutral products were secured by extraction of the alkaline solution by H_2O or petroleum ether.

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"Tartrate Aluminum Compounds," A. V. Pavlinova,
Voronezh State U

"Zhur Obshch Khimii" Vol 17, 1947, pp 3-10

The K compound was prepared by dissolving 100 g $KAl(SO_4)_2 \cdot 12H_2O$ in 80 ml warm water, adding 40 g $K_2C_4H_4O_6 \cdot 2H_2O$ and about 25 g KOH until the solution was slightly alkaline to phenolphthalein at boiling, and adding gradually water to dissolve the precipitate forming on addition of the KOH. The solution was evaporated until a crystal film appeared, cooled, decanted from the K_2SO_4 crystals, and precipitated with an equal volume of alcohol. Full details and chemical formula given.

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